



# Tuneable properties of carbon quantum dots by different synthetic methods

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## Abstract

Carbon quantum dots (CQD) were prepared from three different precursors and by three bottom-up synthesis methods: classical pyrolysis of citric acid (CAP), microwave irradiation of glucose (GM), and hydrothermal treatment of glucosamine hydrochloride (GAH). CQD were further functionalized using various nitrogen-containing compounds: 6-aminohexanoic acid, 1,6-diaminohexane, N-octylamine, dimethylamine, and tryptophan. Special attention was dedicated to investigate how the combination of synthetic method and starting material affected the nature and properties of CQD. The analysis indicated that CAP were good candidates for covalent post-functionalization, GM allowed an easy passivation, and GAH permitted the direct introduction of nitrogen into the core. The size distribution showed a core-shell structure for CQD functionalized with an aminoacid by microwave irradiation, whereas the thermal decomposition evidenced the degradation of functionalizing molecules and the presence of pyridinic and pyrrolic nitrogen after hydrothermal synthesis. Photoluminescence spectra revealed important differences between the synthesis techniques, related to the occurrence of surface states, and the highest fluorescence quantum yield for hydrothermally prepared CQD. These approaches led to CQD with properties that can be exploited in many fields from energy conversion to sensing.

**Keywords** Pyrolysis · Microwave irradiation · Hydrothermal treatment · Photoluminescence · Carbon nanodots

## Introduction

Carbon quantum dots (CQD) are fascinating nanomaterials with tuneable band gap and high photoluminescence (PL) efficiency [1, 2]. The easy synthesis [3], the absence

of toxicity and the huge variability by a vast number of functionalizing molecules [4] are very attractive features that explain the current excitement about CQD.

CQD are semiconducting or insulating particles depending on the nature of the functionalizing groups [5]. For example, nitrogen doping by urea [6] leads to a band gap decrease, because some donor nitrogen atoms are introduced into the graphene planes [7]. Given their semiconducting properties and environmental friendliness, CQD were proposed as noble-metal free catalytic electrodes for the oxygen reduction reaction (ORR) in fuel cells [8]. The emission ranges of photoluminescence can also be tuned by proper functionalization [2, 6, 9, 10]: for example, the blue photoluminescence of unfunctionalized CQD is shifted to green by amines [9, 11].

Various synthetic bottom-up methods were reported in the literature [12–15]. A simple pyrolysis approach was developed by carbonization of common organic molecules [16, 17]. Starting from citric acid (CA), CQD and graphene oxide (GO) can be obtained tuning the carbonization degree of the precursor. Depending on the carbonization temperature or

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time, CA can be incompletely carbonized affording CQD with a strong photoluminescence activity or, after complete pyrolysis, GO with a weak and excitation-dependent PL activity [17, 18]. The application of this method is to some extent challenging, since the structure of CQD cannot be exactly controlled due to carbonaceous aggregation during carbonization [19]. Furthermore, pyrolysis is a time-taking process and a careful monitoring of the reaction is needed.

Microwave irradiation is another effective method to prepare CQD by providing a homogeneous heat distribution to the precursor solution, which quickens the formation of CQD and saves time. Zhu et al. synthesized CQD by treating poly(ethylene glycol) and saccharide solutions for few minutes in a microwave oven [14]. Chandra et al. prepared green fluorescent CQD from sucrose and phosphoric acid in a short time [20]. Advantages of this method include the ability to stop microwave radiation instantly, reducing sample overheating, and the opportunity to use low reaction temperatures, which makes microwave methods reproducible and simple. The elevated pressure that the reaction system can reach during microwave irradiation is the main disadvantage of this technique [21].

The hydrothermal method, generally performed in an autoclave, is nowadays very popular, because the process can be controlled easily. Usually, aqueous solutions with the appropriate precursors are heated directly. Various molecules were used as precursors for CQD, including citric acid and amino acids. The hydrothermal treatment of renewable resources, such as aloe, lychee peel, silk or pepper [22–24], was also proposed to obtain CQD in accordance with “green chemistry”. Due to the high pressure and temperature, the reaction yield is generally high, but the control over CQD size is poor [25].

The functionalization procedures of CQD play a fundamental role not only in the PL modulation, but also in many other applications including ORR catalysis [26–28], selective oxidation of organic molecules [29], doping of semiconducting membranes [30], drug delivery [19, 23], chemical or biosensing [4, 31] etc. The functionalization can lead to bulk doping or to definite defect organizations, such as surface passivation by noncovalent bonds or by covalent linkages of different molecules in edge sites.

In this work, we investigated how the combination of three synthetic approaches, three starting materials, and five N-based modifiers could influence CQD properties. The aim was to find the best conditions to tailor the nature of CQD to achieve nitrogen doping, surface passivation by molecules with Van der Waals interactions, and strong covalently bonded structures.

The synthetic methods included the classical pyrolysis technique, the microwave irradiation route and the hydrothermal approach in an autoclave reactor, while the different green precursors were citric acid, glucose, and glucose

amine hydrochloride. The N-based modifiers ranged from hydrophilic to hydrophobic molecules, such as aminoacids, diamine or octylamine. The tailored CQD were studied in terms of composition, microstructure, thermal stability, and photoluminescence.

## Experimental

### Materials

D-(+)-glucose (G), D-(+)-glucosamine hydrochloride (GA), citric acid (CA), 6-aminohexanoic acid (6AHA), N-octylamine (Oct), N,N-dimethylamine (DMA), 1,6-diaminohexane (DAH), tryptophan (Try), and the other chemicals were reagent-grade and were used as received from Aldrich.

The following abbreviations are used in the text.

CAP: CQD from citric acid by pyrolysis.

CAPDMA: CQD from citric acid functionalized with dimethylamine.

CAPOct: CQD from citric acid functionalized with N-octylamine.

GM: CQD from glucose by microwave irradiation.

GM6AHA: CQDs from glucose functionalized with 6-aminohexanoic acid.

GMTry: CQDs from glucose functionalized with L-tryptophan.

GAH: CQDs from glucosamine hydrochloride by hydrothermal synthesis.

GAHOct: CQDs from glucosamine hydrochloride functionalized with N-octylamine.

GAHDAH: CQDs from glucosamine hydrochloride functionalized with 1,6-diaminohexane.

### Synthesis

#### CQD from citric acid by pyrolysis

**Bare CQD (CAP)** 2 g of citric acid (15 mmol) were inserted in a 3-neck flask equipped with a water condenser and treated in a heating mantle at 200 °C for 10 min. During that time, citric acid melted and the colour of the liquid changed from colourless to yellow and finally to orange. The liquid was cooled to room temperature and its pH was adjusted at 9 adding dropwise a 1.5 M NaOH solution under stirring. The final aqueous solution was filtered (0.2 µm membrane filter) and purified in 2 kDa dialysis bag.

**CQD functionalized with dimethylamine (CAP-DMA)** Using the same apparatus setup, CAPDMA were obtained cooling the orange liquid before solidification and adding DMA (2 M in THF, 7.5 mL) under nitrogen flux (molar reagent ratio CA:DMA = 1:1). The mixture

was kept under stirring and heated at 65 °C for 24 h, then 1.5 M NaOH solution was added dropwise to the viscous mixture at room temperature up to pH = 9. The final mixture was centrifuged (9500 rpm, 30 min) and the solid dissolved in 19 mL of deionized water. This solution was purified as previously reported.

**QCD functionalized with N-octylamine (CAPOct)** Following the procedure used for CAPDMA, QCD functionalized with N-octylamine were obtained adding 5 mL of THF and 2.5 mL of N-octylamine to the orange liquid (molar reagent ratio CA:Oct = 1:1). THF was added to reproduce the experimental condition used for CAPDMA. A sudden change of colour to dark violet was observed. After 24 h at 65 °C the dark greenish THF solution was centrifuged (9500 rpm, 40 min) obtaining a sediment and a liquid part used for the characterization.

#### QCD from glucose by microwave irradiation

**Bare QCD (GM)** 1.0 g of glucose (5.5 mmol), 19 mL of deionized water, and 1 mL of 2 M HCl were put in a glass bottle. The mixture was stirred for about 10–15 min until a clear solution was obtained, then placed in a microwave oven at a power of 800 W for 4 min. After this time, a black precipitate was observed; the glass bottle was cooled to room temperature and the precipitate dissolved under stirring in 10 mL of deionized H<sub>2</sub>O for 5–10 min. The brown solution was centrifuged (9500 rpm, 30 min) obtaining a sediment and a liquid part used for the characterization.

#### QCD functionalized with 6-aminohexanoic acid (GM6AHA)

0.51 g of glucose (2.8 mmol) and 0.37 g of 6-aminohexanoic acid (2.8 mmol) were dissolved under stirring in 19 mL of H<sub>2</sub>O and 1 mL of 2 M HCl in a 500 mL glass bottle. After 30 min, a clear solution was obtained and placed in a microwave oven following the above procedure. The formed black precipitate was dissolved in 10 mL of deionized H<sub>2</sub>O and the solution was stirred for about 5–10 min and then centrifuged. No sedimentation was observed. Elemental analysis (atom%): O 33.4, C 47.4, N 5.2, and H 7.9.

**QCDs functionalized with L-tryptophan (GMTry)** The same procedure was carried out for GMTry. Glucose and L-tryptophan (2.5 mmol, ratio 1:1) dissolved in 19 mL of H<sub>2</sub>O and 1 mL of 2 M HCl were placed in a microwave oven at a power of 8000 W for 4 min. The black precipitate was dispersed in water and centrifuged (9500 rpm, 40 min). After the centrifugation, the sediment part was removed and the

brown liquid portion centrifuged again and then used for the analysis.

#### QCD from glucosamine hydrochloride by hydrothermal synthesis

**Bare QCD (GAH)** 0.45 g of glucosamine hydrochloride (2.1 mmol) was put in 1 mL of H<sub>2</sub>O and stirred for 40–50 min at 50 °C to obtain a solution. The solution was then transferred in an autoclave reactor and inserted into a preheated oven at 190 °C [32]. After 12 h the autoclave was cooled down to room temperature and a black soft solid was recovered and dispersed in deionized water. After centrifugation (9500 rpm, 40 min) a small residue was separated and the solution was used for further characterization.

**QCD functionalized with octylamine (GAHOct)** 0.45 g of glucosamine hydrochloride (2.1 mmol) and 0.27 g of N-octylamine were mixed in 1 mL of H<sub>2</sub>O and stirred for 1 h at 50 °C. The solution was put in an autoclave as reported before. The formed black solid was insoluble in water; chloroform was used for the dispersion. The centrifugation (8500 rpm, 10 min) afforded a liquid part, used for the characterization, and a small residue.

**QCD functionalized with 1,6-diaminohexane (GAH-DAH)** The same procedure was carried out for the GAH-DAH sample, using 0.45 g of GA and 0.24 g of 1,6-diaminohexane (molar ratio 1:1). After the autoclave reaction the solid was dispersed in methanol and then centrifuged (8500 rpm, 10 min). No sedimentation was observed.

#### Characterization

##### FTIR

FTIR spectra were recorded in transmission mode in the range of 4000–400 cm<sup>-1</sup> using a spectrometer Perkin-Elmer Spectrum 100. The dry samples were analysed in cesium iodide cells. Spectra were normalized against a background spectrum.

##### NMR

<sup>13</sup>C NMR spectra were recorded with a Bruker AVANCE III spectrometer operating at 400 MHz. The samples were dried and then dissolved in the deuterated solvent (DMSO d<sub>6</sub>, D<sub>2</sub>O, and toluene d<sub>8</sub>). Chemical shifts (ppm) were referenced to tetramethylsilane (TMS).



## Transmission Electron Microscopy (TEM)

TEM samples were prepared from a droplet dried in carbon film under infrared radiation for a maximum duration of 120 s. The carbon film had a mean thickness of 50 nm. All TEM inspections were carried out with a Philips<sup>TM</sup> CM-20<sup>®</sup> working at 200 kV with a double-tilt specimen holder equipped with liquid nitrogen cooling stage. To acquire a sound statistical evaluation, a minimum of three samples of CAP and GM6AHA were inspected.

## X-ray photoelectron spectroscopy (XPS)

The samples for X-ray photoemission spectroscopy analyses were prepared by depositing a drop of solution on grated Au foil (purity 99.99%), then they were dried in air. Photoemission spectra were acquired using a spectrometer Escalab 250Xi (Thermo Fisher Scientific Ltd., East Grinstead, UK) with a monochromatic Al K $\alpha$  (1486.6 eV) excitation source and six-channeltron detection system. All the spectra were collected at a pass energy of 40 eV in standard mode of electromagnetic lens system corresponding to an analysis area of approximately 1 mm in diameter. The binding energy (BE) scale was corrected for a low static charging by positioning the C 1 s peak of aliphatic carbon (C–C, C–H bonds) at BE = 285.0 eV and controlling if the Fermi level corresponds to BE = 0 eV. Spectroscopic data were processed by the Avantage v.5 software (Thermo Fisher Scientific Ltd).

## Thermogravimetric analysis (TGA)

The thermograms were recorded using a high resolution thermogravimetric analyser TA Q500. The experiments were conducted with Pt sample holders, a CQD mass of about 7 mg, a temperature range 50–700 °C and a nominal heating rate of 3 K/min under air flux.

## Optical measurements

All the measurements for optical characterization were performed by keeping the solutions of CQD in fused-silica cuvettes with an optical path-length of 10 mm. UV–Vis absorption spectra were recorded using a Varian Cary 50 Scan spectrophotometer. Photoluminescence (PL) emission spectra were acquired with a specific laboratory setup based on a continuous-wave 200-W Hg(Xe) discharge lamp (Oriol Corp.), an excitation 25-cm monochromator (PMI), an emission 25-cm monochromator (Oriol Cornerstone 260) and a Hamamatsu R3896 photomultiplier for detection of the emission light. Appropriate optical filters and the conventional 90° geometry were used for improved stray-light rejection and signal-to-noise ratio. PL measurements were performed in diluted solutions in spectroscopic grade solvents,

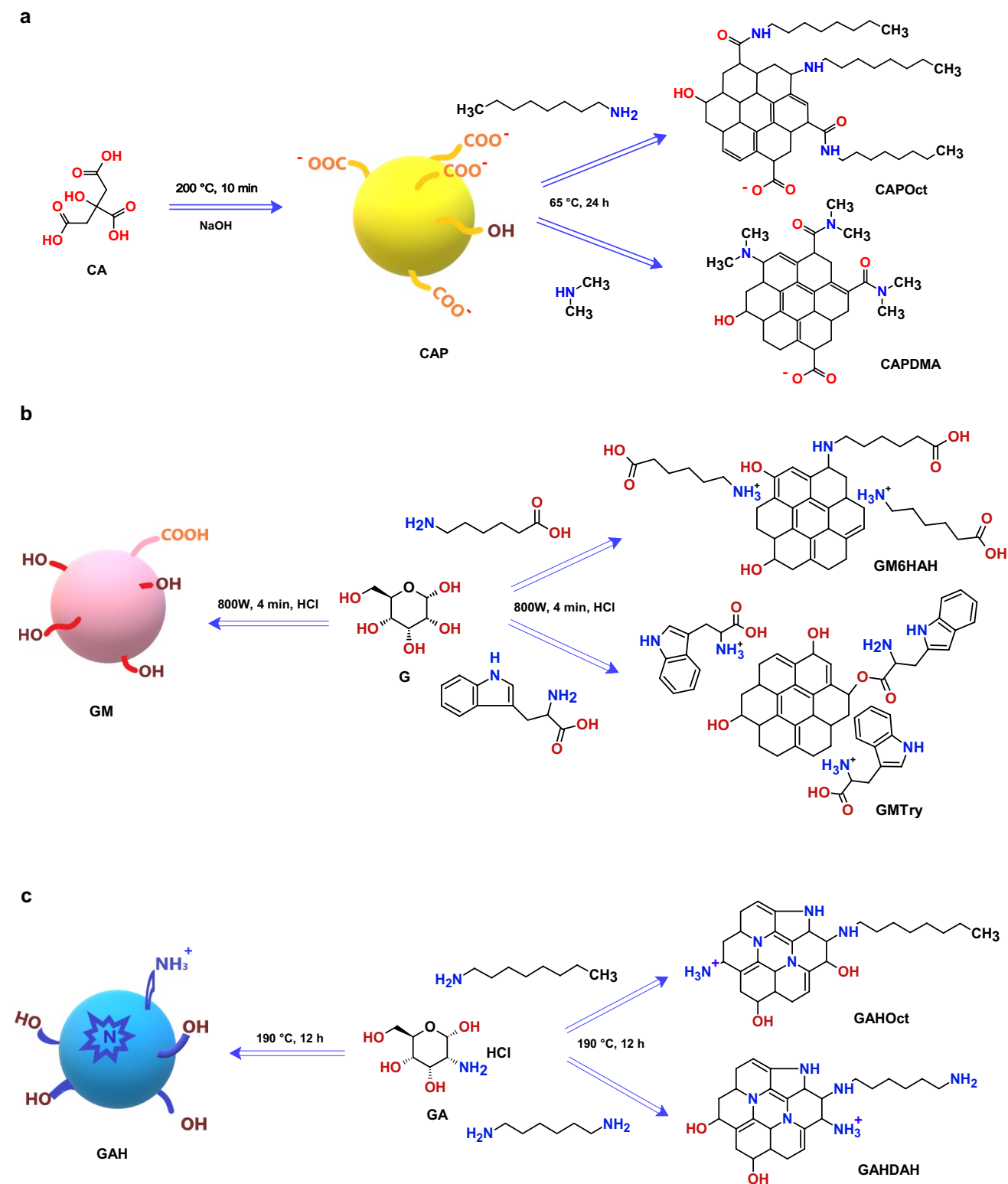
with an absorbance < 0.1 optical density (OD) in the whole excitation/emission range. The typical excitation wavelength of  $\lambda_{\text{exc}} = 355$  nm was used with a spectral bandpass of 3 nm for both the excitation and emission monochromators. The spectra were corrected for the instrumental response through a calibration performed using a reference black-body lamp and standard-fluorophore solutions. The relative quantum yield of the fluorescent emission was determined using both quinine sulphate and 9,10-Diphenylanthracene as reference and solutions with an absorbance < 0.05 OD in the whole excitation/emission range.

## Results and discussion

Scheme 1 shows the paths followed for the synthesis of bare and functionalized CQD. The structure of the products proposed in the Scheme is in agreement with the results obtained from the different analysis. The synthesis by pyrolysis is reported to start with the thermal decomposition of CA [33]. CA dehydrates around 175 °C with the formation of aconitic acid, the decomposition continues until 225 °C reaching 80% of weight loss around 200 °C. This temperature, selected in the current work, affords small particles, around 1.4 nm diameter, although the PL, after a treatment of 10 min, is relatively low [34]. The time of 10 min at 200 °C was chosen to have a balance between the hydroxyl and carboxyl groups, used for the subsequent functionalization, and the graphitic part of CQD. A longer time would lead to an increase in carbon–carbon bonds to the detriment of the oxygenated groups [34]. The formation of CQD is considered to take place via the synthesis of aromatic clusters by aldol condensation and cycloaddition, reactions catalysed typically in basic media. The balance of pH can be critical to achieve the aggregation reactions between carbon nanoparticles and an optimum value of pH = 9 was chosen for obtaining CQD with small size [3]. The bare CAP were post-functionalized with N,N-dimethylamine and N-octylamine. Due to the low reaction temperature (65 °C), the functionalization occurs via the formation of amide and amine bonds (see FTIR discussion) and nitrogen is not part of the skeletal network (Scheme 1a).

Both the synthesis of GM and GAH begin from carbohydrates. These starting materials are very attractive due to their accessibility, high solubility in water, and low carbonization temperatures [35]. The reaction starts by the opening of the anomeric carbon followed by polymerisations and aromatic or hetero-aromatic cluster formation [36, 37].

The microwave mediated glucose reaction is an acid-catalysed dehydration that produces alcohols and carboxylic acids on the CQD surface responsible of the solubility in polar solvents (Table 1). The surface passivation was obtained by reaction of glucose in the presence of amino



**Scheme 1.** Schematic synthesis routes for CQD made by (a) pyrolysis, (b) microwave irradiation, (c) hydrothermal treatment

acids. As reported in Scheme 1b, nitrogen is linked as amine to the CQD shell or present as passivating agent (see FTIR, NMR, and TEM). This behaviour was attributed to the weak

interaction between glucose and functionalizing compounds by mostly van der Waals forces and hydrogen bonds [38]. It was also reported that CQD synthesized by microwave





**Table 1** Solubilities of CQD in various solvents. + : >5 g/L

Solvent and Dielectric constant ( $\epsilon$ )	H <sub>2</sub> O	DMSO	DMAc	DMF	MeOH	NMP	THF	CHCl <sub>3</sub>	Toluene
Samples									
CAP	+	+	-	+	+	-	-	+	-
CAPDMA	+	+	-	+	-	-	-	+	-
CAPOct	-	-	+	+	-	+	+	+	+
GM	+	+	+	+	-	+	-	-	-
GMTry	+	+	+	+	+	+	+	-	-
GM6AHA	+	+	+	+	+	+	-	-	-
GAHOct	-	+	+	+	+	+	-	+	+
GAHDAH	+	+	+	+	+	+	-	-	-

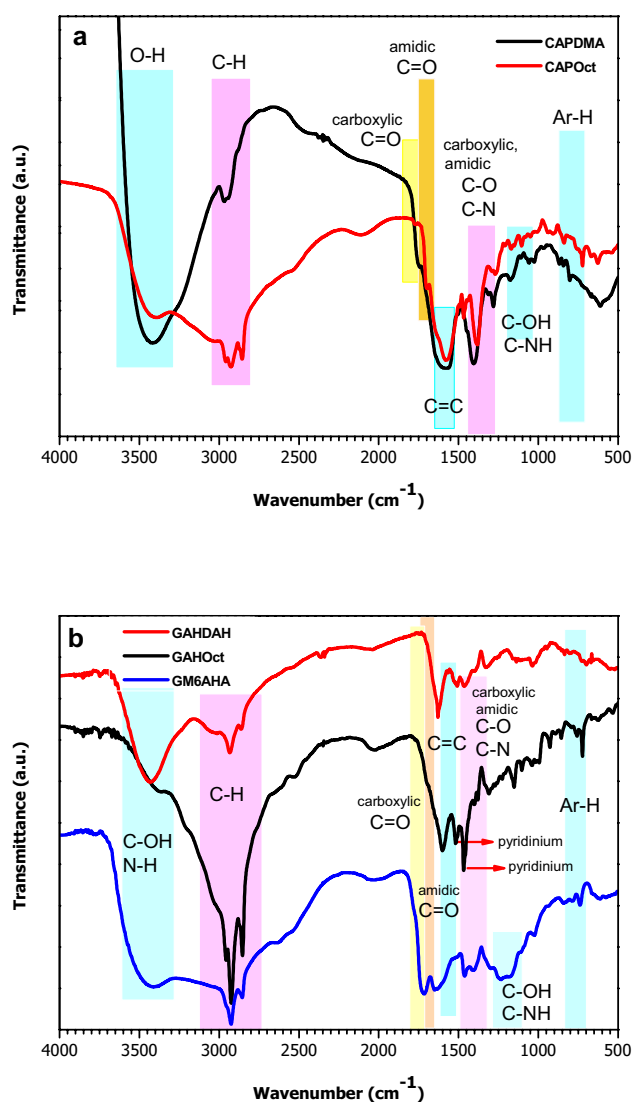
irradiation in the presence of water present a self-passivated layer on the surface [39].

CQD prepared from glucosamine hydrochloride by hydrothermal synthesis offer the advantage to contain nitrogen in the starting material. The reported mechanism includes the oligomerization of glucosamine molecules by intermolecular dehydration followed by polymerization and aromatization via intramolecular dehydration [40]. The functionalization was realized using 1,6-diaminohexane and N-octylamine and resulted in an increase of graphitic nitrogen.

The colloidal solubilities reported in Table 1 show that only CQD functionalized with octyl groups are insoluble in water, due to the hydrophobic nature of the long aliphatic chain. In fact, the CQD-Oct are the only samples soluble in toluene confirming the lyophilic nature. The CQD synthesized by microwave irradiation are instead insoluble in solvents with low dielectric constant and show a larger hydrophilicity, in agreement with the core-shell structure evidenced by TEM and XPS (see below). CQD by pyrolysis present a variable solubility without a clear trend in function of the dielectric constant indicating a more complicated substituent distribution.

Figure 1 illustrates the FTIR analysis of functionalized samples. FTIR spectra of bare samples and GMTry are reported in Supplementary Information (SI 1).

OH stretching is observed for all compounds around 3450 cm<sup>-1</sup>. The intensity is lower for the octylamine-functionalized CQD (Fig. 1a: CAPOct and Fig. 1b: GAHOct) in agreement with the hydrophobic nature of the long aliphatic chain. The C-H stretch of alkyl groups at 2960 cm<sup>-1</sup> is evident for all compounds, but less pronounced for CAPDMA due to the absence of a long chain existing in all other samples. The CO stretching of carboxylic acid, originating from CA precursor, is present for CAPDMA at 1750 cm<sup>-1</sup> and less evident for CAPOct. The amide CO vibration (1680 cm<sup>-1</sup>), showing a reaction with the amine group of the functionalizing molecule, is strong for CAPOct, while it exists as shoulder in CAPDMA. Regarding GM and GAH samples,



**Fig. 1** Comparison of FTIR spectra for CQD made by: **a** pyrolysis (CAPDMA and CAPOct); **b** microwave irradiation (GM6AHA) and hydrothermal synthesis (GAHOct and GAHDAH)

the CO stretching of carboxylic acid is at  $1750\text{ cm}^{-1}$  for the bifunctional GM6AHA, while the CO amide vibration is present as a shoulder for GAHDAH and GAHOOct in the range  $1660\text{--}1680\text{ cm}^{-1}$  and seems absent in GM6AHA sample because of the non-covalent core-shell structure. At  $1640\text{ cm}^{-1}$ , the bending absorption of N-H is marked for GAHDAH, due to the two amine groups of its functionalized molecule; this band is less evident for GAH6AHA. The aromatic C=C vibration around  $1610\text{ cm}^{-1}$  is observed especially for CAPDMA, CAPOct, and GAHOOct, less intense for GM6AHA and GAHDAH. The peak at  $1530\text{ cm}^{-1}$ , strong for GAHOOct and weak for GAHDAH, can be assigned to the bending vibration of pyridinium ions [41]; this peak is associated with the bending of pyridine around  $1450\text{ cm}^{-1}$ , visible in both samples of Fig. 1b, and as shoulder in CAPDMA (Fig. 1a). CAPDMA and GM6AHA show an absorption at  $1400\text{ cm}^{-1}$  characteristic of C-O vibrations of carboxylic acid and [42]. In the other samples, this signal is less pronounced or shifted (CAPOct,  $1360\text{ cm}^{-1}$ ) indicating the transformation of carboxylic acid into amide, in agreement with the previous consideration. A band at  $1315\text{ cm}^{-1}$  is observed for GAHOOct and can be assigned to graphitic nitrogen, in agreement with XPS results [43]. At  $1150\text{ cm}^{-1}$  the C-N stretching is observed especially for GAHOOct, GADAH, and CAPDMA. In the range  $700\text{--}760\text{ cm}^{-1}$  aromatic C-H-bending vibrations are visible. A summary of peak intensities and assignments is reported in SI 2.

$^{13}\text{C}$  NMR spectra of various samples are reported in Fig. 2. Three main regions are roughly identified in the spectrum of Fig. 2a. The  $\text{sp}^3$  zone can be further subdivided in high-field and low-field regions, where alkylic and C-X (X = O, N) carbons absorb.

The CAPDMA spectrum shows that the nitrogen of the DMA moiety is connected to CQD by a direct linkage (N-CH<sub>3</sub> and C-N peaks around 45 and 55 ppm) and an amidic bond (CH<sub>3</sub>-N-CO 35-38 ppm and N-C=O around 170 ppm) as represented in Scheme 1a. In the region of  $\text{sp}^2$  carbons, the peaks between 120 and 130 ppm are ascribed to graphitic C, such as polycyclic aromatic carbons, while peaks from 130 to 140 ppm can be assigned to internal C=C bonds. The existence of peaks between 110 and 120 ppm was attributed to the presence of terminal C=C bonds [44-46].

Peaks of amide moieties can be recognized around 170 ppm and around 180 ppm. The CAPOct spectrum (Fig. 2b) does not show substantial changes from the previous one, except in the aliphatic region, where octylamine absorbs. For this sample, it is possible to conclude that the amine is linked to CQDs by aminic and amidic linkages, in agreement with the FTIR results. The spectrum is recorded in toluene to favour a good solubility of this hydrophobic sample, although the solvent obscures the  $\text{sp}^2$  region. More complex is the GM6AHA spectrum

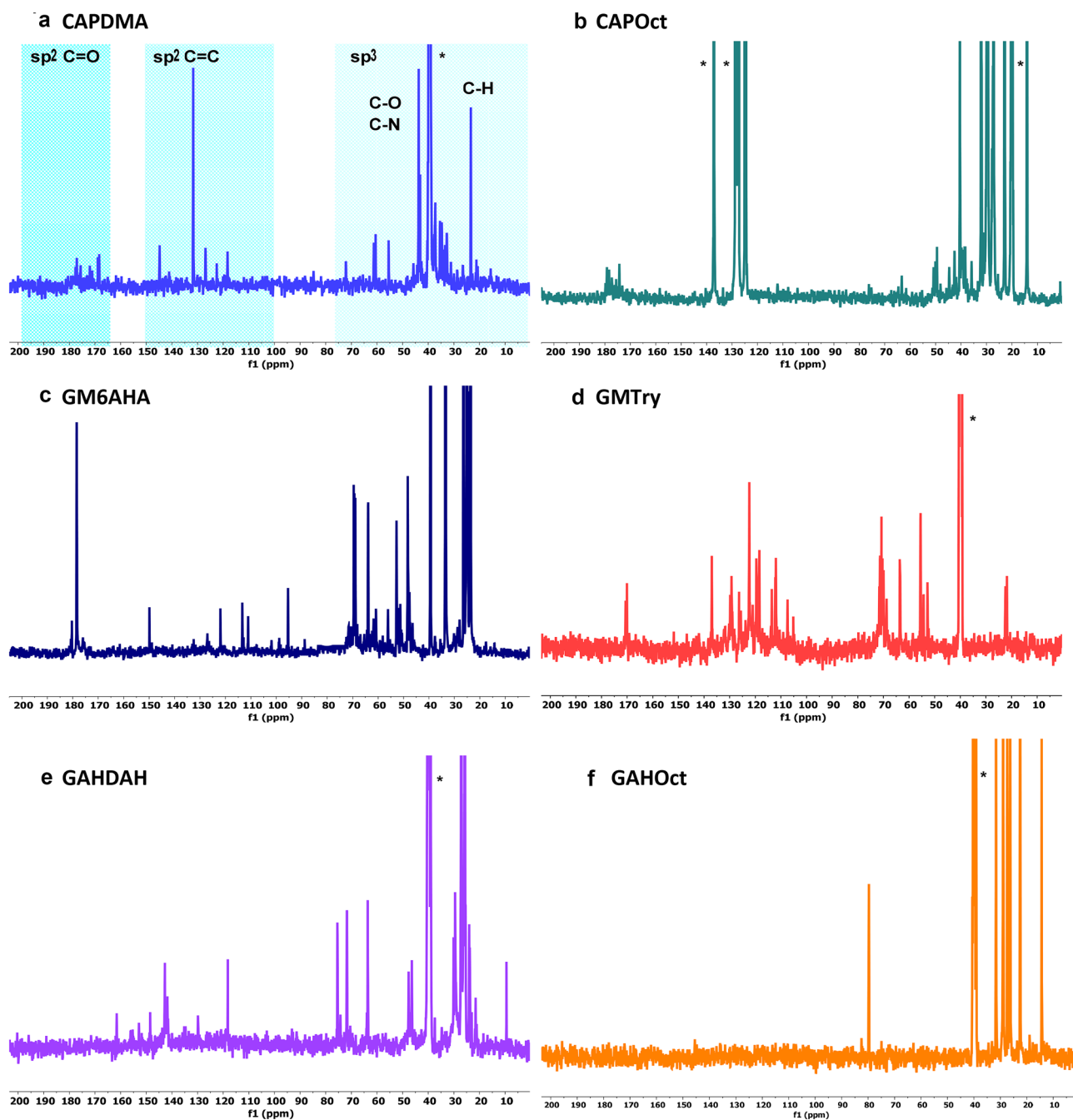
(Fig. 2c) especially in the range 40-70 ppm, where there are the absorbances of C-O and C-N carbons, indicating a considerable amount of alcoholic and aminic groups also deriving from the functionalizing molecule. The peak at 95 ppm can be ascribed to isolated double bonds in enolate form, and the absorbance at 150 ppm to phenol or aniline moieties. This hypothesis is strengthened by the low quantity, or absence, of the amidic peaks around 170 ppm, confirming the FTIR observations. Figure 2d shows the NMR spectrum of GMTry. In the range 108-137 ppm, it is possible to recognize the absorbance of tryptophan that maintains its structure after the synthesis of CQD. These peaks obscure the absorption of graphitic and polycyclic carbons. The tryptophan original signals at 56 and 175 ppm ( $\alpha$  carbon and carboxylic acid bearing in  $\alpha$  an amino group, respectively) appear shifted in the spectrum (Fig. 2d) suggesting that a linkage with CQD occurs via ester and/or amino bonds.

The spectra of CQDs synthesized by the hydrothermal method present a different pattern and no carboxylic groups are present in Fig. 2e, f. Peaks around 150-160 ppm in GAHDAH (Fig. 2e) can be assigned to graphitic nitrogen-carbon bonds together with peaks at 140 ppm indicating graphitic carbons. Signals around 120 ppm can be attributed to pyrrolic groups, while in the range 60-70 ppm, we can recognize C-O and C-N peaks. Solubility issues affect the quality of the last spectrum, reported in Fig. 2f. The aliphatic part is very similar to that reported in Fig. 2b. A very small peak can be detected in the region 140-150 ppm, corresponding to C=C bonds.

The TEM micrographs of CAP (Fig. 3a) and GAH (Fig. 3b) show bare CQD with a size distribution around an average of about 4 nm, in good agreement with the literature [47]. Instead, the CQD functionalized with 6-aminohexanoic acid (GM6AHA, Fig. 3c, d) presents a core-shell structure. The shell width of around 0.9 nm is in perfect agreement with the six-carbon chain molecule. The core size distribution around an average value of 2.4 nm corresponds to around 16 carbon atoms in each direction, which is equivalent to about 6 six-membered rings or around 200 carbon atoms in the CQD core. Similar core-shell structures, confirmed by SAXS patterns, were reported recently with amino-acid-like molecules. The shell around CQD appears to be a weakly bonded reaction intermediate that collapses in later reaction stages or with higher energy methods, such as pyrolysis and hydrothermal synthesis [48].

Figure 4 shows XPS spectra of the N 1s region including the peak fitting for the investigated samples and the survey spectrum for GMTry. Table 2 presents the BE values in eV and atomic concentrations of different nitrogen species. The spectra of the C 1s region for the samples GAHOOct, GAH, GMTry, and CAP are reported in the SI section (SI 3-7) with survey spectra and the attributions of carbon species.





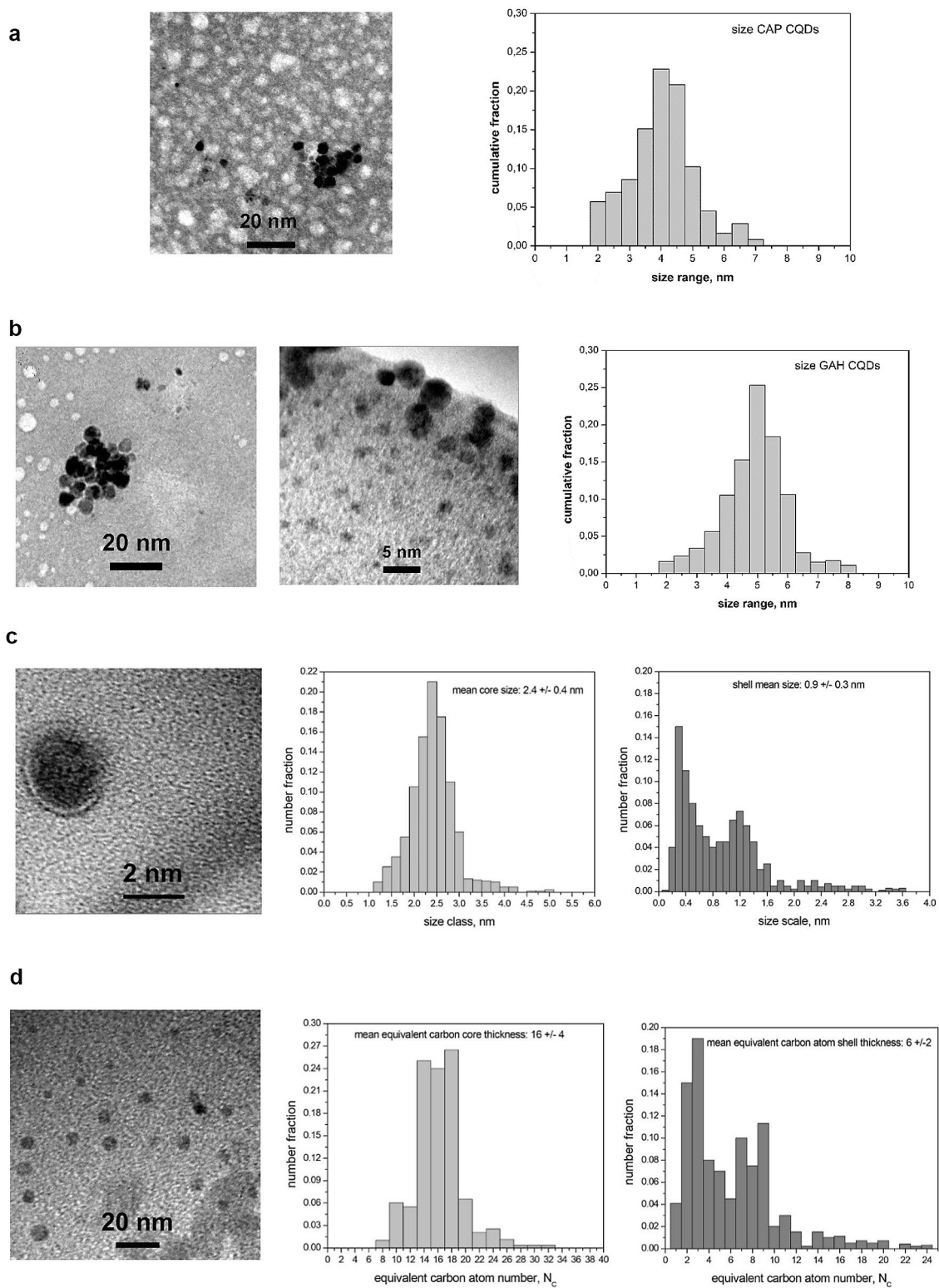
**Fig. 2.**  $^{13}\text{C}$  NMR spectra of: (a) CAPDMA ( $\text{DMSO } d_6$ ); (b) CAPOct (toluene  $d_8$ ); (c) GM6AHA ( $\text{D}_2\text{O}$ ); (d) GMTry ( $\text{DMSO } d_6$ ); (e) GAHDAH ( $\text{DMSO } d_6$ ); (f) GAHOct ( $\text{DMSO } d_6$ )

As described in ref. [43], it is possible to identify the chemical state of nitrogen in the CQD structure from XPS results: the N 1s peaks for graphitic, pyrrolic, and pyridinic nitrogen in N-doped graphene are at BE of 401.5, 399.7, and 397.9 eV, respectively, whereas the hydrogenation of pyridinic N shifts its peak to 400.5 eV. Slightly different values for the N 1s peak have been reported in ref. [7]: graphitic N at 401.3 eV, pyrrolic N at 400.2 eV and pyridinic N at

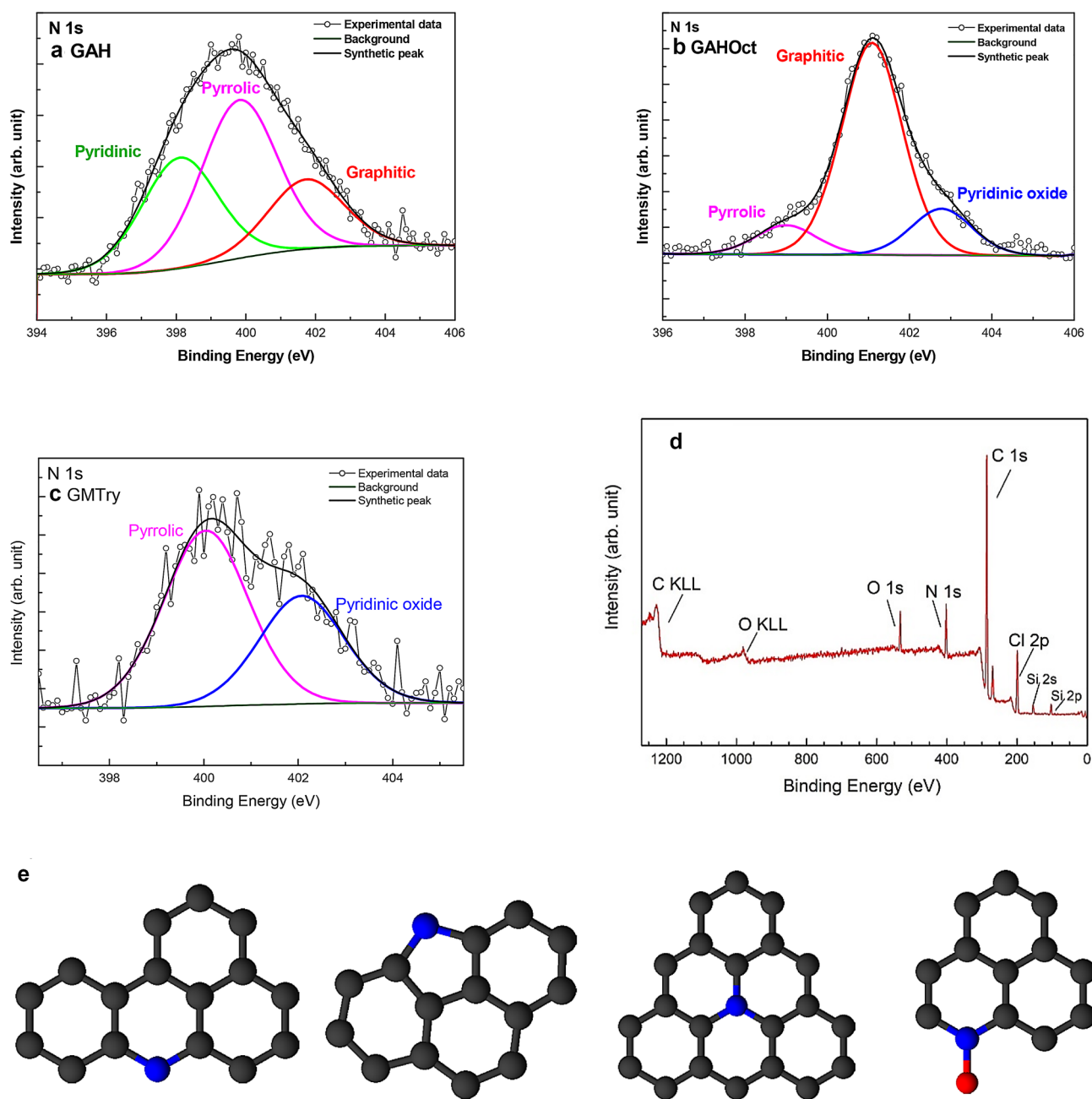
398.5 eV. More attributions of N 1s peaks can be found in the refs. [49–52].

Pyrrolic nitrogen (BE  $\approx$  400 eV) is observed in all the samples. The XPS results also indicate that during the hydrothermal synthesis some nitrogen is inserted in graphitic position (BE  $\approx$  401 eV), especially when octylamine is present. The pyridinic nitrogen (BE  $\approx$  398 eV) is observed only in the GAH sample that was prepared without functionalizing





**Fig. 3** BF-TEM micrographs and CQD size distribution for (a) CAP, (b) GAH and (c), (d) GM6AHA



**Fig. 4** XPS spectra of the N 1s region for three samples: a) GAH, b) GAHOct c) GMTry; d) Survey spectrum of GMTry. The Si signals originate from the sample holder; e) Schematic representation of vari-

ous nitrogen bond types (from the left: pyridinic, pyrrolic, graphitic, pyridinic oxide)

**Table 2** BE values and atomic concentrations of various nitrogen species in CQD

Nitrogen species	Pyridinic	Pyrrolic/Amine/Amide	Graphitic	Pyridinic oxide
Sample				
GAH	398.0 (2.5 at%)	399.7 (3.7 at%)	401.6 (1.6 at%)	–
GAHOct	–	399.0 (1 at%)	401.1 (7.4 at%)	402.8 (1.6 at%)
GMTry	–	400.1 (3.1 at%)	–	402.0 (2.3 at%)

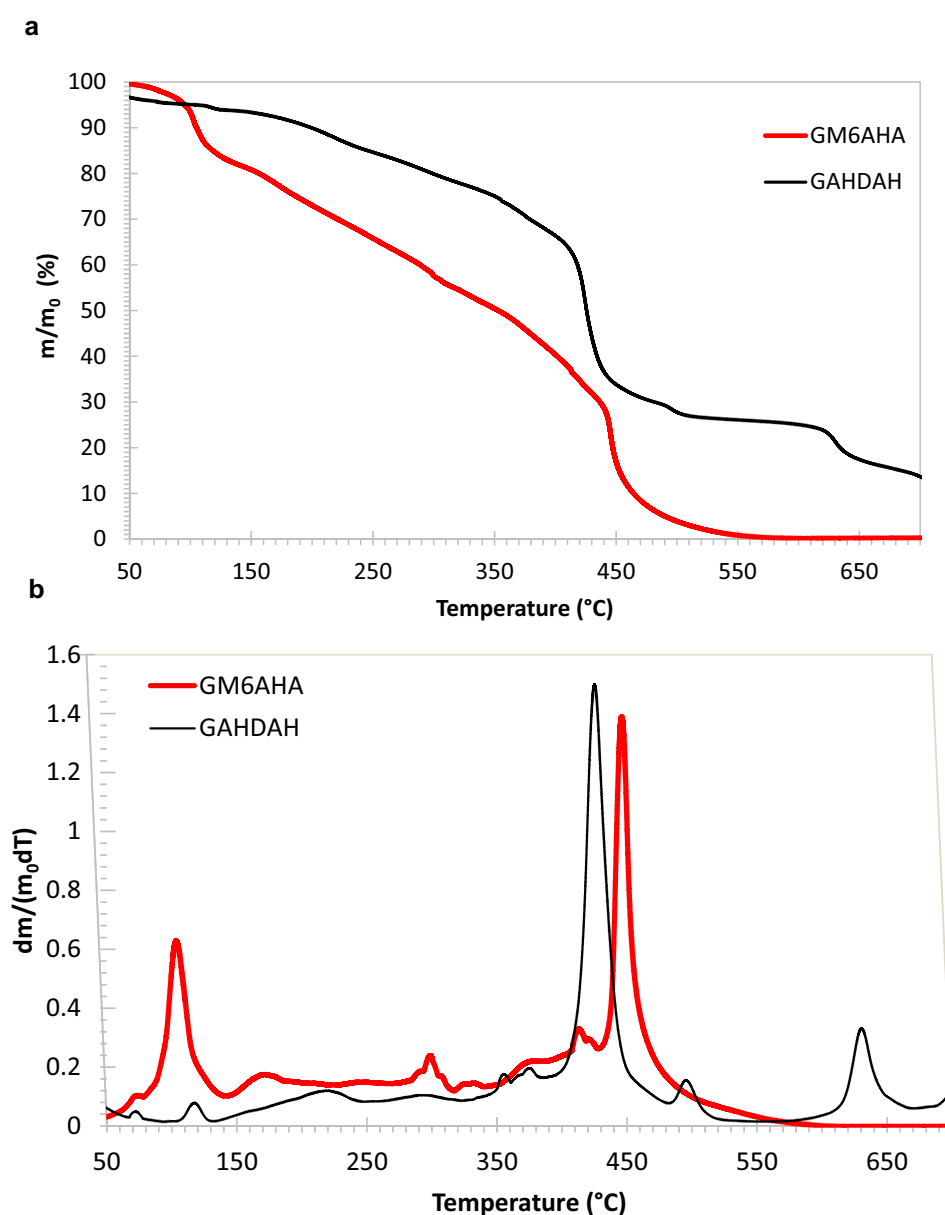


molecule in hydrothermal conditions, whereas the pyridinic oxide species ( $BE \approx 402$  eV) are only visible, when functionalizing molecules were added. The co-existence of pyrrolic and pyridinic oxide species and the absence of graphitic nitrogen in the sample GMTry are consistent with the presence of the functionalized molecules in a shell around the CQD core, such as in the case of GM6AHA observed by TEM. Both samples made by microwave synthesis present thus a core-shell structure, which is related to the lower energy provided by this method.

Figure 5 shows thermogravimetric curves (mass loss and derivative) for two samples made by hydrothermal synthesis and microwave irradiation. The peaks in the derivative curves correspond to the various thermal degradation processes. A first mass loss with a maximum around  $100$  °C

is related to the loss of chemisorbed water. [29, 53] It is much more pronounced in the case of GM6AHA due to the hydrophilic nature of the carboxylic acid groups of 6-aminohexanoic acid, which strongly increases the hydrophilicity of this sample that is water soluble. This observation is in agreement with the core-shell structure of GM6AHA evidenced by TEM indicating that the functionalizing molecules are intact. Only a small water loss is noticed in the sample made by hydrothermal synthesis, because most of the diamino-hexane has been transformed into structural nitrogen in the CQD. A small peak at  $75$  °C is probably related to physisorbed water on less hydrophilic amine and oxygenated groups, present in both samples. The broad mass loss between  $150$  and  $400$  °C in GAHDAH and GM6AHA is typical of CQD and is generally attributed to the loss of

**Fig. 5** Thermogravimetric analysis of CQD from glucose functionalized with 6-aminohexanoic acid by microwave irradiation (GM6AHA) and from glucosamine hydrochloride functionalized with 1,6-diaminohexane by hydrothermal synthesis (GAHDAH): **a** mass loss curve, **b** derivative curve



oxygenated groups, including hydroxyl, ketone, carboxylic acid, and amide [54]. The larger mass loss for GM6AHA indicates a larger amount of oxygenated groups [29] in agreement with elemental analysis, NMR, and XPS. The main decomposition process of CQD is observed with a peak maximum at 425 (GAHDAH) and 450 °C (GM6AHA), corresponding to the high stability of aromatic bonds in CQD.

The presence of the core–shell structure seems to enhance the thermal stability of GM6AHA. Small peaks about 40 K below the main peak are observable in both samples and correspond probably to a minor amount of  $sp^3$  carbon, as observed previously in poly(2,6-dimethyl-1,4-phenylene oxide), where the structural methyl groups are degraded before the aromatic main chain [55]. One notices that the degradation of the core–shell GM6AHA is terminated at 550 °C, whereas a plateau is observed in the case of GAHDAH. This plateau was previously attributed to pyridinic and pyrrolic nitrogen, which present a high thermal stability and decompose only around 625 °C [54].

Figure 6 displays the UV–Vis absorption spectra of the reference compounds, as prepared by the different methods of synthesis and starting materials without any functionalization. The absorbance spectrum of CAP (black continuous line) shows the typical profile found in the literature for this type of bottom-up preparation [17, 56]. The general steep decrease of absorbance with increasing wavelength is usually assigned to the tail of high-energy  $\pi^*-\pi^*$  transitions of aromatic C=C bonds [17, 56]. Differently, a controversial attribution has been given to the smooth peak/shoulder at 340 nm. In the first work on CA-derived CQDs [17], the peak was supposed to originate from optical transitions in quantum-size ordered clusters of  $sp^2$  carbon atoms confined within a defective  $sp^3$  matrix. More recently, also in the framework of the results reported in several types of CQDs

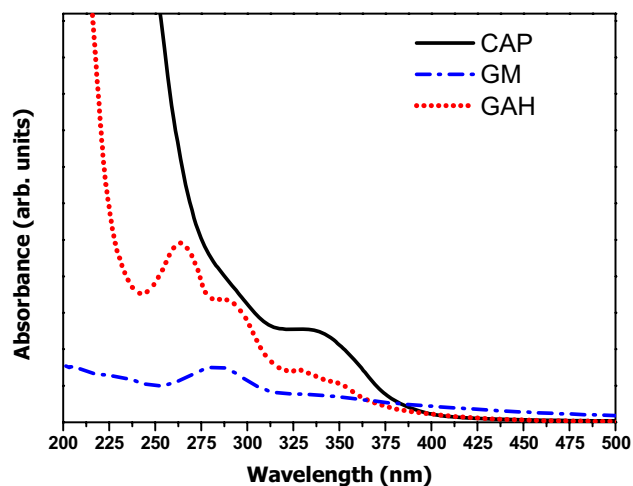
prepared by top-down synthesis [57–59], similar structures were attributed to the  $n-\pi^*$  transitions of surface states [56]. Surface states are synergistic hybridizations of the carbon backbone and linked oxygen-containing functional groups, such as carboxyl, hydroxyls and epoxy [57]. In the present case, a much smoother shoulder is also visible at 295 nm and is possibly originating from surface states, as well. In GM (blue dash/dot line), the structure at 340 nm is almost missing, while the shoulder at 295 nm has become a small yet clearly visible peak slightly shifted at 290 nm, also found in previous reports on microwave-assisted synthesis [60]. Differently, when glucosamine and hydrothermal synthesis are used (GAH, red dotted line), the curve profile becomes more complex, with a distinct third peak at 270 nm that adds up to the smooth structures at 295 nm and 340 nm.

The differences between pyrolysis and microwave syntheses in the relative intensities of the structures at 290 and 340 nm could originate from a diverse distribution and composition of the oxygen-containing functional groups. On the other hand, the structure at 270 nm arising in GAH is located at energies appreciably closer to those of  $\pi^*-\pi^*$  transitions of aromatic C=C bonds. This observation suggests that this peak can be assigned to new optical transitions due to nitrogen doping originating from the starting material and hydrothermal synthesis [56, 61]. In fact, the presence of graphitic and pyridinic nitrogen is confirmed by XPS data (see Table 2).

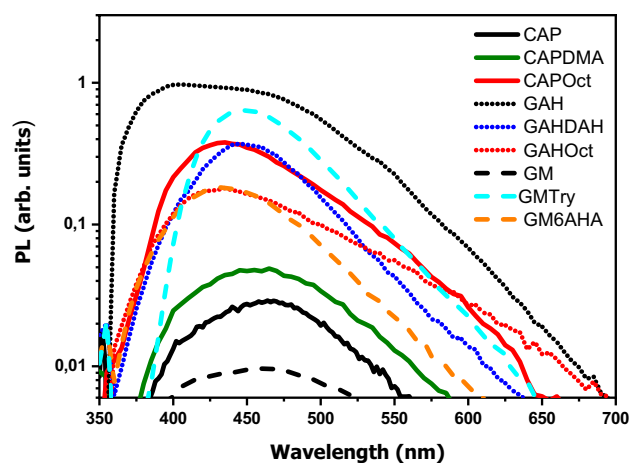
The UV–Vis absorption spectra of CAP derivatives are reported in SI 8. The data show that the intensity of the spectral structures at 290 and 340 nm progressively decreases with functionalization, especially in CAPOct. This can be explained by the nitrogen moieties substituting the oxygen-containing functional groups (Scheme 1), thus lowering the related absorbance [61].

Fig. SI 9 displays the UV–Vis absorption spectra of GM derivatives. In this series, functionalization with 6-amino-hexanoic acid in GM6AHA only increases and slightly shifts the peak at 290 nm of the reference compound. Conversely, functionalization with tryptophan in GMTry produces a peculiar spectrum with a remarkable band at 280 nm, very similar to the absorption feature of pure tryptophan itself, and a smaller peak at 360 nm. Similar results have also been observed in functionalized CQD prepared with different methods [62], thus suggesting that this amino-acid retains its electronic level configuration when it is bound to CQD.

Finally, as regards the hydrothermal series (Fig. SI 10), functionalization with 1,6-diaminohexane in GAHDAH does not change the spectrum of the reference compound significantly. On the other hand, the curve of GAHOct is remarkably flat and does not present any appreciable structure, confirming the strong effect of octylamine on the functional groups, as observed in CAPOct, and its capacity to insert



**Fig. 6** UV–Vis absorption spectra of CQD made by: pyrolysis (CAP); microwave irradiation (GM); hydrothermal treatment (GAH)



**Fig. 7** PL emission spectra of CQD made by the different methods and various functionalization

nitrogen in graphitic position, as confirmed by XPS data (see Table 2).

In summary, if CAP are taken as reference compound/method, the greatest and most significant changes of UV–Vis absorbance are produced using nitrogen-containing starting materials, especially octylamine, combined with hydrothermal synthesis. Smaller modifications are obtained with the less energetic microwave methods, except for functionalization with tryptophan that gives a material, which remarkably retains the absorption characteristics of the pure fluorophore along with a peculiar lower energy peak.

Figure 7 shows the photoluminescence (PL) emission spectra of the various CQD along with the respective derivatives and different functionalization. A semi-log plot has been used to due to large differences in the relative intensities and to display them more clearly. The excitation wavelength was  $\lambda_{\text{ex}} = 355 \text{ nm}$  for all the samples. The spectra are normalized by the optical density at excitation wavelength, so that the plotted curves are representative of the relative fluorescence quantum yield.

As can be seen, all samples exhibit the blue fluorescence in the range around 450 nm which is typical of CA-derived CQD [17]. An appreciable short-wavelength contribution, which shifts the peak to 400 nm, is only observed in GAH, while a smaller blue shift to 430 nm is produced by functionalization with octylamine and 6-aminohexanoic acid. However, it is evident that the functionalization does not alter the profile of emission spectra as much as it does for the absorption curves. This suggests that neither linked amino groups nor nitrogen doping of carbon backbone appreciably modify the recombination centers for excitons that are the surface defects due to non-perfect  $\text{sp}^2$  domains [56]. Instead, very different of PL signal can be observed, with variations well above one order of magnitude clearly indicating that functionalization affects the fluorescence quantum yield  $\Phi$  of

**Table 3** Quantum yield for the different CQD

Sample	$\Phi$ (%)
CAP	$0.6 \pm 0.2$
CAPDMA	$1 \pm 0.2$
CAPOct	$7 \pm 0.2$
GAH	$24 \pm 2$
GAHDAH	$6 \pm 0.2$
GAHOct	$4 \pm 0.5$
GM	$0.2 \pm 0.1$
GMTry	$10 \pm 2$
GM6AHA	$4 \pm 0.5$

the emission process, as reported in Table 3. Especially, QY decreases in the order:  $\text{GAH} > \text{GMTry} > \text{CAPOct} > \text{GAHDAH} > \text{GM6AHA} \cong \text{GAHOct} > \text{CAPDMA} > \text{CAP} > \text{GM}$ . We note that higher values of QY are obtained when nitrogen-containing starting materials or functionalization are used. This is in agreement with the hypothesis that the introduction of nitrogen can produce new surface states that trap electrons and increase radiative recombination [56, 61]. In particular, the highest QY is observed in GAH that shows the presence of graphitic, pyrrolic and pyridinic nitrogen (see Table 2). Graphitic nitrogen can enhance the rigidity of the CQD structure thus reducing non-radiative relaxation through lattice vibrations [56].

Finally, Table 4 shows a comparison of the three synthesis methods.

## Conclusion

We studied the synthesis of CQD starting from three “eco-friendly” molecules, citric acid, glucose and glucoseamine, by three different methods, pyrolysis, microwave irradiation, and hydrothermal synthesis. The CQD were functionalized using various nitrogen-containing compounds.

The investigation by various spectroscopic and thermal analysis methods and observation by transmission electron microscopy shows that the nature of CQD reflects the chemistry of the starting materials combined with the synthetic method.

The pyrolysis of citric acid leads to CQD rich of carboxylic moieties, very versatile groups that can be easily transformed into amide groups by reaction with amines. The nature of the functionalizing molecule influences the CQD hydrophilic/hydrophobic behaviour. The microwave irradiation of glucose gives materials with mainly alcoholic groups; the reaction in the presence of amino acids affords passivated CQD. This effect is probably related to the positive charge on the amino acids in acidic condition that allows strong electrostatic interactions. The hydrothermal treatment of glucosamine provides CQD with nitrogen in





**Table 4** Main characteristics of each synthesis

	Mechanism	Advantages	Disadvantages	Main functional groups	Features	Applications	Time/Cost
Pyrolysis (CAP)	Aldol Condensation and cycloaddition, Basic catalysis	Uniform particle size (pH control), Ease of scale-up, tunable carbonization degree	Poor control over structure due to carbonaceous aggregation, Low product yield	Carboxylic, some hydroxyl, graphitic carbon	Post-functionalization of surfaces	Polymer/CQD (photo-voltaics [63], supercapacitors [64])	Medium/low
Microwave (GM)	Acid-catalysed dehydration, polymerization, Aromatic clusters, carbonization	Homogeneous, reproducible and simple, clean	Small volumes, High pressure of the system, energy consumption, low product yield	Hydroxyl, some carboxylic	Short-chain passivated surfaces	Electroluminescent light emitting diodes [65], bio imaging [21]	Short/medium
Hydrothermal (GAH)	Inter- and intramolecular dehydration, heteroaromatic clusters, polymerization	Easy control, High product yield, eco-friendly	Poor size control	N-doped skeleton, Amine, hydroxyl, no carboxylic groups	High fluorescence quantum yield	Electrocatalysis [8], sensing [66], Batteries [67], CO <sub>2</sub> transformation [68]	Long/low

graphitic, pyrrolic and pyridinic positions. Further reactions with amines do not alter the CQD skeleton, but enhance the amount of nitrogen in the material.

The properties and the possible applications of CQD can thus be tuned: pyrolyzed CQD are functionalized by covalently bonded molecules and can be used as synthon compounds for example in the interaction with polymers. CQD obtained by the less energetic microwave process present passivation and formation of a core-shell structure with essentially Van der Waals bonds, very interesting for bio-applications. Finally, CQD by hydrothermal processes, the most energetic ones, exhibit bulk insertion of nitrogen that can act as a donor dopant. These samples show by far the highest quantum yields fundamental for sensing applications.

**Supplementary Information** The online version contains supplementary material available at <https://doi.org/10.1007/s40097-021-00431-8>.

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**Availability of data and materials** Not applicable.

**Code availability** Not applicable.

**Declaration**

**Conflict of interest** No conflict of interest.

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